When magnesium metal burns in air, it combines with oxygen to form magnesium oxide.

Write the BCE for the above reaction with the proper subscripts.

_____

Confirm this with someone!

What type of reaction is this: _____

A) How many moles of oxygen are required to produce 10.0 moles of magnesium oxide?

B) How many moles of magnesium oxide are produced by the reaction of 130. g of magnesium?

C) If 2.60 g of oxygen react, how many moles of magnesium oxide will form?

D) What mass of oxygen combines with 10.00 g of magnesium in this reaction?

E) What mass of magnesium oxide will be produced by the reaction of 45.00 g of magnesium?

F) If 3.6×10^{26} molecules of oxygen are to react, how many moles of magnesium are required?

G) If 1.50 g of magnesium reacts, how many oxygen atoms are involved?