Chapter 6 pg 224-245

What you should know about gases. Nelson Chemistry References

*-kinetic molecular theory pg 233 + notes ⇒ explains all!

-manometers (open and closed) and barometers

-Dalton's Law of Partial Pressures (not in text)

-Boyle's Law pg 225

-Charles' Law pg 230

-Celsius scale ⇒ never use for gas problems!!

-Kelvin scale + 273 pg 229

-absolute zero 0K = −273 °C = Charles' law lab with syringes + water baths

-Gay-Lussac's Law pg 234

-Diver's Law

-Combined gas law pg 231

-ideal versus real gases

-ideal gas equation pg 243

-density of a gas

-molar mass of a gas = butane + large ∉

-molar volume of a gas pg 236

*-STP = 0°C and 101.3 kPa

-molar volume of a gas at STP $Mg + HCl \rightarrow H_2$ gas Lab

-stoichiometric problems = mass to volume; volume to volume; volume to mole

-collection of a gas over water

- SATP = standard ambient T & P = 25°C & 100 kPa
- Graham's Law (not in text)