The Trend for Ionization Energy for the First 20 elements of the Periodic Table

IA								VIIA	VIIIA
H 1312.0	IIA			IIIA	IVA	۷A	VIA	H 1312.0	He
Li	Be			В	C	N	0	F	Ne
520.2	899.4			800.6	1086.4	1420.3	1313.9	1681.0	2080.6
Na	Mg			Al	Si	P	S	Cl	Ar
495.8	737.7	-Carrer	200000	577.6	786.4	1011.7	999.6	1251.1	1520.5
K	Ca			Ga	Ge	As	Se	В	Kr
418.8	589.8	100	-	578.8	762.1	947	940.9	1139.9	1360.7
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On a separate sheet of graph paper:

- 1) Write down the definition of Ionization Energy.
- 2) Draw a graph of Ionization Energy vs Atomic Number--go big or go home--≥ 75 % of the page!!
- 3) Describe what happens as you go across a period.

Explain why this happens.

4) Describe what happens as you go down a group.

Explain why this happens.