

The Trend for Ionization Energy for the First 20 elements of the Periodic Table

IA						VIIA	VIIIA
H 1312.0						H 1312.0	He 2372.3
Li 520.2	Be 899.4					F 1681.0	Ne 2080.6
Na 495.8	Mg 737.7					Cl 1251.1	Ar 1520.5
K 418.8	Ca 589.8					B 1139.9	Kr 1380.7
		III A	IV A	V A	VI A		
		B 800.6	C 1086.4	N 1420.3	O 1313.9		
		Al 577.6	Si 786.4	P 1011.7	S 999.6		
		Ga 578.8	Ge 762.1	As 947	Se 940.9		

On a separate sheet of graph paper:

- 1) Write down the definition of Ionization Energy.
- 2) Draw a graph of Ionization Energy vs Atomic Number--go big or go home--> 75 % of the page!!
- 3) Describe what happens as you go across a period.
Explain why this happens.
- 4) Describe what happens as you go down a group.
Explain why this happens.