**The Ka of Boric Acid**

**Purpose:**

* to determine the Ka of a weak acid, H3BO3

**Materials:**

* 2 burets--Read to 2 Decimal Places
* Erlenmeyer flask
* White paper
* BTB or PHTH
* Funnel
* Beaker of NaOH
* Boric Acid
* Various pH papers

**Procedure:**

* determine the pH of the boric acid
* titrate the boric acid with NaOH

**Observations:**

**Calculations:**

**BCE titration:**

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**1) concentration of H+ ion**

**2) concentration of the boric acid**

**3) Ka calculation**

**Conclusion:**

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Why is Boric Acid considered a weak acid? You must comment on equilibrium. Explain and justify.

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