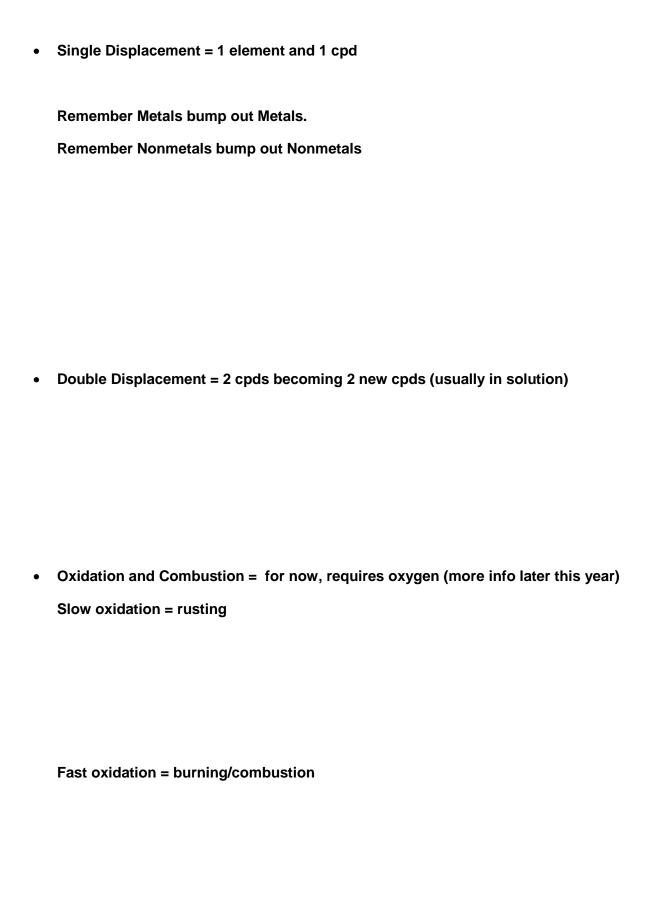
Stuff to know when writing Balanced Chemical Equations (BCE)

| 1. | | | s are placed first in a compound (ionic cpd). Always use charges when s are involved. |
|----|---|----------------|--------------------------------------------------------------------------------------------------------------------------|
| 2. | | Use s | ubscripts. |
| | • | you c | an determine whether a precipitate might form using your solubility chart |
| | | e.g. | Nal(aq) + Pb(NO ₃) ₂ (aq) → |
| | • | Reme | ember all group 1A and NH ₄ ⁺ and NO ₃ ⁻ cpds are soluble in water i.e. (aq) |
| | | e.g. | NaCl(aq) |
| | | o.g. | Pb(NO ₃) ₂ (aq) |
| | | | |
| | | | (NH ₄) ₂ CO ₃ (aq) |
| | • | you s condi | hould know the state of certain chemicals at room temperature and pressure tions |
| | | e.g. | H ₂ O(I) |
| | | | Cl ₂ (g) |
| | | | Al(s) |
| | | | |
| | | _ | |
| 2. | | P | S I H ave N o Br ight O r Cl ever F riends |

| 3. | Be able to determine the type of chemical reaction. | |
|------------|--------------------------------------------------------------------------|---------|
| J . | | J L - |
| | Refer to pages 25 - 27 in your textbook. Make notes for yourself if need | a be |
| | Types: | |
| • | Acid + Base = Salt + Water | |
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| | | |
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| • | Synthesis = a cpd forming from elements or cpds | |
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| | | |
| | | |
| | Decembesition — a and breaking apart into elements and/or simpler on | مام |
| • | Decomposition = a cpd breaking apart into elements and/or simpler cp | us |
| | | |
| | | |
| | | |
| | | |
| | | |
| | Precipitation | |
| | | |
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| • | Photosynthesis and Respiration |
|------|-----------------------------------------------------------------------------|
| | |
| | |
| | |
| • | Endothermic and Exothermic |
| | Endothermic = More energy is absorbed during the reaction than is released. |
| | |
| | Exothermic = More energy is released than is absorbed during the reaction. |
| Fire | Triangle |
| | |