**Stuff to know when writing BCE**

**1. Metals are placed first in a compound (ionic cpd). Always use charges when metals are involved.**

**2. Use subscripts.**

* **you can determine whether a precipitate might form using your solubility chart**

**e.g. NaI(aq) + Pb(NO3)2(aq) → \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

* **Remember all group 1A and NH4+ and NO3- cpds are soluble in water i.e. (aq)**

**e.g. NaCl(aq)**

**Pb(NO3)2(aq)**

**(NH4)2CO3(aq)**

* **you should know the state of certain chemicals at room temperature and pressure conditions**

**e.g. H2O(l)**

**Cl2(g)**

**Al(s)**

**2.**  **P S I H ave N o Br ight O r Cl ever F riends**

**3. Be able to determine the type of chemical reaction.**

**Refer to pages 25 - 27 in your textbook. Make notes for yourself if need be.**

**Types:**

* **Acid + Base = Salt + Water**
* **Synthesis = a cpd forming from elements or cpds**
* **Decomposition = a cpd breaking apart into elements and/or simpler cpds**
* **Precipitation**
* **Single Displacement = 1 element and 1 cpd**

**Remember Metals bump out Metals.**

**Remember Nonmetals bump out Nonmetals**

* **Double Displacement = 2 cpds becoming 2 new cpds (usually in solution)**
* **Oxidation and Combustion = for now, requires oxygen (more info later this year)**

**Slow oxidation = rusting**

**Fast oxidation = burning/combustion**

* **Photosynthesis and Respiration**
* **Endothermic and Exothermic**

**Endothermic = More energy is absorbed during the reaction than is released.**

**Exothermic = More energy is released than is absorbed during the reaction.**

**Fire Triangle**