

## TABLE of COMMON IONS

## COMMON POSITIVE IONS (CATIONS)

aluminium	$\text{Al}^{3+}$	lead(IV)	$\text{Pb}^{4+}$
ammonium	$\text{NH}_4^+$	lithium	$\text{Li}^+$
barium	$\text{Ba}^{2+}$	magnesium	$\text{Mg}^{2+}$
calcium	$\text{Ca}^{2+}$	manganese (II)	$\text{Mn}^{2+}$
caesium	$\text{Cs}^+$	mercury (I)	$\text{Hg}_2^{2+}$
chromium (II)	$\text{Cr}^{2+}$	mercury (II)	$\text{Hg}^{2+}$
chromium (III)	$\text{Cr}^{3+}$	nickel(II)	$\text{Ni}^{2+}$
copper (I)	$\text{Cu}^+$	potassium	$\text{K}^+$
copper (II)	$\text{Cu}^{2+}$	silver	$\text{Ag}^+$
hydrogen	$\text{H}^+$	sodium	$\text{Na}^+$
iron (II)	$\text{Fe}^{2+}$	tin (II)	$\text{Sn}^{2+}$
iron (III)	$\text{Fe}^{3+}$	tin (IV)	$\text{Sn}^{4+}$
lead (II)	$\text{Pb}^{2+}$	zinc	$\text{Zn}^{2+}$

## COMMON NEGATIVE IONS (ANIONS)

acetate	$\text{CH}_3\text{COO}^-$	nitride	$\text{N}^{3-}$
bromide	$\text{Br}^-$	nitrite	$\text{NO}_2^-$
carbonate	$\text{CO}_3^{2-}$	oxalate	$\text{C}_2\text{O}_4^{2-}$
hydrogen carbonate	$\text{HCO}_3^-$	oxide	$\text{O}^{2-}$
chlorate	$\text{ClO}_3^-$	permanganate	$\text{MnO}_4^-$
perchlorate	$\text{ClO}_4^-$	phosphide	$\text{P}^{3-}$
chloride	$\text{Cl}^-$	phosphate	$\text{PO}_4^{3-}$
chlorite		(mono)hydrogen phosphate	

## COMMON POSITIVE IONS (CATIONS)

	$\text{ClO}_2^-$		$\text{HPO}_4^{2-}$
hypochlorite	$\text{ClO}^-$	dihydrogen phosphate	$\text{H}_2\text{PO}_4^-$
chromate	$\text{CrO}_4^{2-}$	sulphate	$\text{SO}_4^{2-}$
dichromate	$\text{Cr}_2\text{O}_7^{2-}$	hydrogen sulphate	$\text{HSO}_4^-$
cyanide	$\text{CN}^-$	sulphide	$\text{S}^{2-}$
fluoride	$\text{F}^-$	hydrogen sulphide	$\text{HS}^-$
hydride	$\text{H}^-$	sulphite	$\text{SO}_3^{2-}$
hydroxide	$\text{OH}^-$	hydrogen sulphite	$\text{HSO}_3^-$
iodide	$\text{I}^-$	thiocyanate	$\text{SCN}^-$
nitrate	$\text{NO}_3^-$	thiosulphate	$\text{S}_2\text{O}_3^{2-}$

NOTE: chromium and manganese have positive elemental ions but negative polyatomic ions.

HISTORICAL NOTE: Anions with *hydrogen* in the name, eg *hydrogen carbonate*, used to be named with a *bi-* as in *bicarbonate*.