

## 27

## NET IONIC EQUATIONS

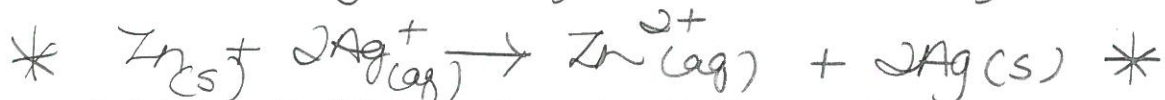
Write nonionic, complete ionic, and net ionic equations to represent the following reactions.

1. A zinc strip is dipped into an aqueous solution of silver nitrate.
2. Solutions of lead(II) nitrate and potassium chloride are mixed.
3. Barium hydroxide is used to neutralize sulfuric acid.
4. Chlorine gas is bubbled through an aqueous solution of sodium iodide.
5. Aqueous solutions of aluminum sulfate and calcium hydroxide are mixed.
6. An iron nail is dropped into a beaker of sulfuric acid.

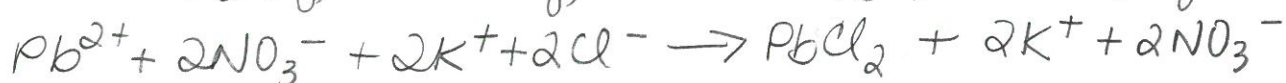
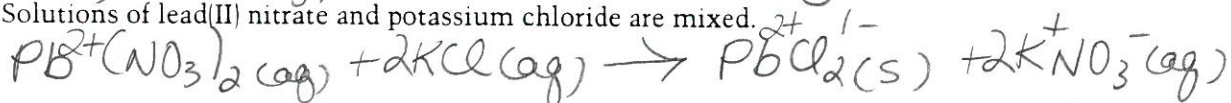
# 27 NET IONIC EQUATIONS

Write nonionic, complete ionic, and net ionic equations to represent the following reactions.

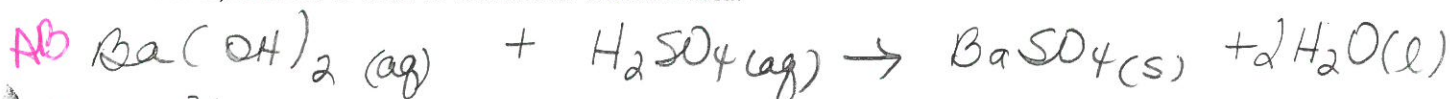
1. A zinc strip is dipped into an aqueous solution of silver nitrate.



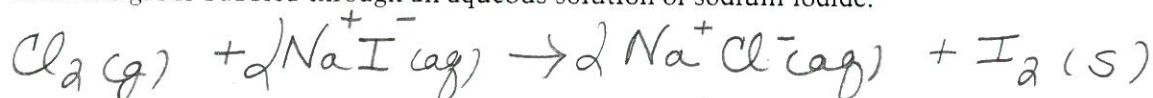
2. Solutions of lead(II) nitrate and potassium chloride are mixed.



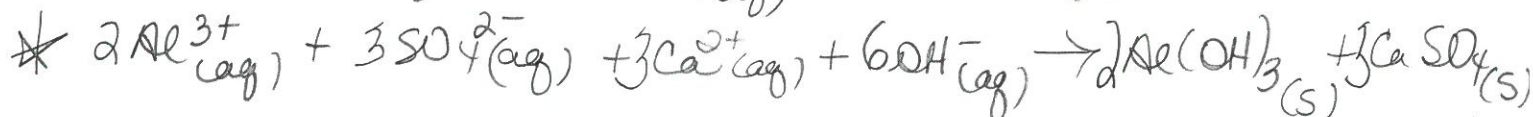
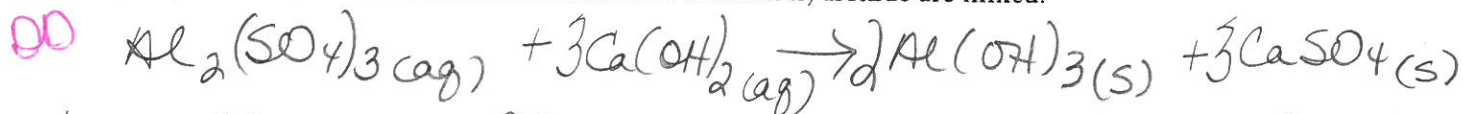
3. Barium hydroxide is used to neutralize sulfuric acid.



4. Chlorine gas is bubbled through an aqueous solution of sodium iodide.



5. Aqueous solutions of aluminum sulfate and calcium hydroxide are mixed.



6. An iron nail is dropped into a beaker of sulfuric acid.

