## NET IONIC EQUATIONS

Write nonionic, complete ionic, and net ionic equations to represent the following reactions.

1. A zinc strip is dipped into an aqueous solution of silver nitrate.

2. Solutions of lead(II) nitrate and potassium chloride are mixed.

3. Barium hydroxide is used to neutralize sulfuric acid.

4. Chlorine gas is bubbled through an aqueous solution of sodium iodide.

5. Aqueous solutions of aluminum sulfate and calcium hydroxide are mixed.

6. An iron nail is dropped into a beaker of sulfuric acid.



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1. A zinc strip is dipped into an aqueous solution of silver nitrate.

$$Z_{n}(s) + \partial Ag^{+}(NO_{3})_{2}(ag) \rightarrow Z_{n}(NO_{3})_{2}(ag) + \partial Ag(s)$$
  
 $Z_{n} + \partial Ag^{+} + \partial NO_{3}^{-} \rightarrow Z_{n}^{2} + \partial NO_{3}^{-} + \partial Ag(s)$   
 $Z_{n} + \partial Ag^{+} \rightarrow Z_{n}(ag) + \partial Ag(s) + \partial Ag(s)$ 

\* The following of lead(II) nitrate and potassium chloride are mixed. He had (S) + 2KNO3 (ag) + 2KCl (ag)  $\rightarrow$  PbCl<sub>2</sub>(S) + 2KNO3 (ag) P62++2N03-+2K++2Cl-->P6Cl2+2K+2N03-

3. Barium hydroxide is used to neutralize sulfuric acid.

Ab Ba (OH) 2 (ag) + H2 SO4 (ag) 
$$\rightarrow$$
 Ba SO4(s) + 2H2O(l)  
\* Ba<sup>2+</sup> + 2OH- + 2H+ + SO42- > Ba SO4(s) + 2H2O(l) \*  
(ag) (ag)

4. Chlorine gas is bubbled through an aqueous solution of sodium iodide.

O 
$$Cl_{a}(g) + dNaT(ag) \rightarrow dNa^{\dagger}Cl_{ag} + I_{a}(s)$$
  
 $Cl_{a} + \partial Na^{\dagger} + \partial I^{-} \rightarrow \partial Na^{\dagger} + \partial Cl^{-} + I_{a}$ 

#  $Cl_{2}(g) + \partial Icag) \rightarrow dCl_{cag}) + I_{2}(5) # 5.$  Aqueous solutions of aluminum sulfate and calcium hydroxide are mixed.

6. An iron nail is dropped into a beaker of sulfuric acid.

Fe(s) + H2504(ag) 
$$\rightarrow$$
 Fe 504 (ag) + H2(g)  
The Fe(s) + 2H (ag) + 504(ag)  $\rightarrow$  Fe (ag) +504(ag) +H2(g)