

TABLE of COMMON IONS

COMMON POSITIVE IONS (CATIONS)

aluminium	Al^{3+}	lead(IV)	Pb^{4+}
ammonium	NH_4^+	lithium	Li^+
barium	Ba^{2+}	magnesium	Mg^{2+}
calcium	Ca^{2+}	manganese (II)	Mn^{2+}
caesium	Cs^+	mercury (I)	Hg_2^{2+}
chromium (II)	Cr^{2+}	mercury (II)	Hg^{2+}
chromium (III)	Cr^{3+}	nickel(II)	Ni^{2+}
copper (I)	Cu^+	potassium	K^+
copper (II)	Cu^{2+}	silver	Ag^+
hydrogen	H^+	sodium	Na^+
iron (II)	Fe^{2+}	tin (II)	Sn^{2+}
iron (III)	Fe^{3+}	tin (IV)	Sn^{4+}
lead (II)	Pb^{2+}	zinc	Zn^{2+}

COMMON NEGATIVE IONS (ANIONS)

acetate	CH_3COO^-	nitride	N^{3-}
bromide	Br^-	nitrite	NO_2^-
carbonate	CO_3^{2-}	oxalate	$C_2O_4^{2-}$
hydrogen carbonate	HCO_3^-	oxide	O^{2-}
bicarbonate	HCO_3^-	permanganate	MnO_4^-
chlorate	ClO_3^-	phosphide	P^{3-}
perchlorate	ClO_4^-	phosphate	PO_4^{3-}
chloride	Cl^-	(mono)hydrogen phosphate	
chlorite	ClO_2^-		

	ClO_2^-		HPO_4^{2-}
hypochlorite	ClO^-	dihydrogen phosphate	$H_2PO_4^-$
chromate	CrO_4^{2-}	sulfate	SO_4^{2-} <i>sulfate</i>
dichromate	$Cr_2O_7^{2-}$	hydrogen sulphate	HSO_4^-
cyanide	CN^- <i>KCN</i>	sulphide	S^{2-}
fluoride	F^-	hydrogen sulphide	HS^-
hydride	H^-	sulphite	SO_3^{2-}
hydroxide	OH^-	hydrogen sulphite	HSO_3^-
iodide	I^-	thiocyanate	SCN^-
nitrate	NO_3^-	thiosulphate	$S_2O_3^{2-}$

NOTE: chromium and manganese have positive elemental ions but negative polyatomic ions.

HISTORICAL NOTE: Anions with hydrogen in the name, eg hydrogen carbonate, used to be named with a bi- as in bicarbonate.