

## Chem Test #1

- 2 question from "The Right Chemistry: Chlorine can save lives or take lives" article. ✓
- Nomenclature (naming chemical compounds; ionic and covalent compounds, don't forget to memorize the highlighted polyatomic ions! Also show the **charges** when showing your work but they don't have to be on the final answer). ✓ x2
- Be able to write an english sentence explaining how a chemical compound comes to be (there is an example on the third page of the "IUPAC Nomenclature or Name That Compound (cpd) -Part 1" package). ✓ ug
- On Solubility chart remember that all Group 1A,  $\text{NH}_4^+$  and  $\text{NO}_3^-$  compounds are all aqueous/soluble in water and know how to use the solubility chart. ✓
- Be able to complete/determine the BCE (don't forget the subscripts!) You will be given an equation and asked to complete it. ✓
- Look over the precipitation lab
- Determine the type of chemical equation (Double displacement, synthesis, combustion etc...).
- Stoichiometry (mole hill and mole box, remember STP-1 mole of any gas will occupy 22.4 L at STP) be able to do stoichiometric calculations. ✓
- Go over both the "IUPAC Nomenclature or Name That Compound (cpd) -Part 1" package and the "Stuff to know when writing BCE" package for additional information! ✓

• limiting reagent

all types of BCE