**Equilibrium Problems involving RICE Tables**

**A(g) + B(g) 🡪 2 C(g)**

|  |  |  |  |
| --- | --- | --- | --- |
| **R**(eaction) | **A** | **B** | **2 C** |
| **I**(nitial concentrations)  You can add anything here in any ratio you want. | Mol or Mol/L | Mol or Mol/L | Mol or Mol/L |
| **C**(hange in concentration) i.e. how much reacted or formed  **Mole ratios from the BCE must be obeyed here!!** |  |  |  |
| **E**(quilibrium concentrations) i.e. the concentrations after the reaction has reached equilibrium. | Mol/L | Mol/L | MolL |

**Overview for Solving Problems Using the Equilibrium Constant**

1. Write the BCE in the R line of the RICE table.

2. Cross out any solids or liquids.

3. Fill in any initial concentrations in the I line. You may use moles or mol/L here.

Pay attention to the VOLUME of the container.

4. Fill in the C line with any changes in the concentrations.

You must obey the mole ratios from the BCE in this line!!!

5. Fill in the E line with the concentrations present at equilibrium.

6. Write the Keq expression using the BCE. No solids or liquids!