

Endothermic and exothermic reactions

EST
PAGES 114-117
Complete this concept review handout and keep it as a record of what you have learned.

EST Definitions

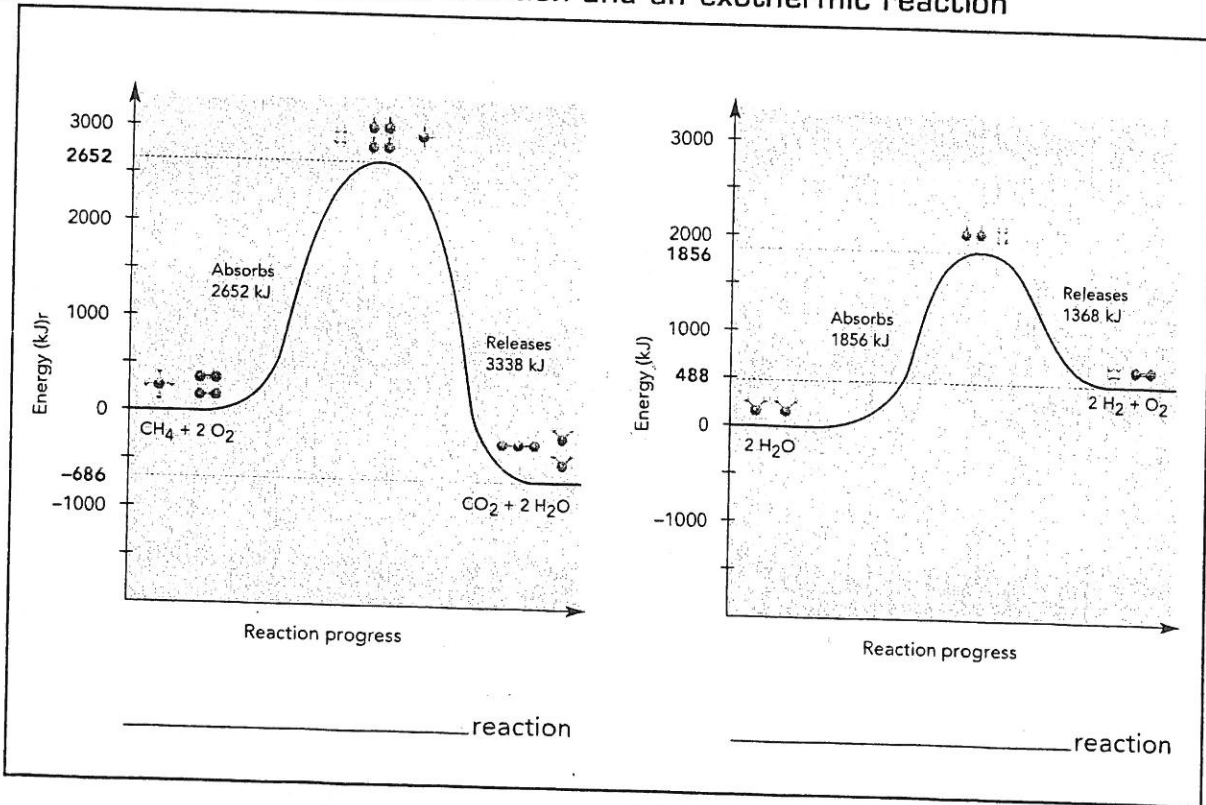
• An exothermic reaction is _____

• An endothermic reaction is _____

EST Reaction energy

The amount of energy released or absorbed by a reaction can be estimated by calculating its reaction energy, which is _____

EST Graphs of an endothermic reaction and an exothermic reaction

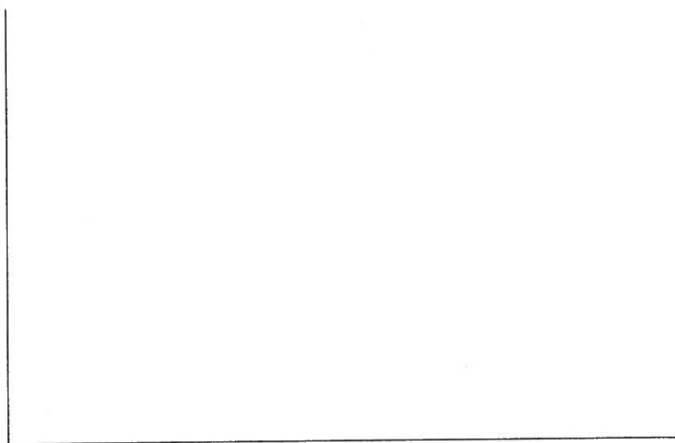


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Problem: How much energy is involved when 1 mole of methane, CH_4 , is burned?

- 1) Write the balanced chemical equation.

- 2) Sketch the Energy graph:



- 3) How many grams of methane will involve 5.7×10^4 kJ of energy?

- 4) If 400 kJ of heat energy are released, how many grams of carbon dioxide will be produced?