

Compounds as ions

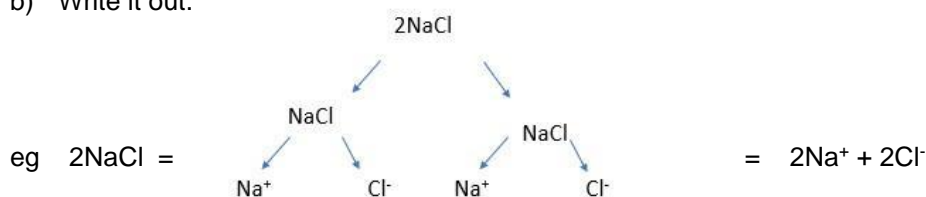
Education in Chemistry

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Split each compound into its ions:

- Draw a graphical representation.
- Write it out.



Section 1:

NaCl
NaF
LiF
KBr
CaO
BaO
BaS
KI
AlN
GaP

Section 2:

2NaBr
3KI
4MgO
InN
5SrS

Section 3:

CaCl₂
AlCl₃
Na₂O
CaO
Al₂O₃
Li₂S
2Li₂S
3Li₂S
5MgCl₂
8Na₃N
Potassium fluoride
Potassium oxide
Potassium nitride
Calcium nitride

Section 4:

NaOH
Na₂SO₄
Na₂CO₃
2Na₂CO₃
6Li₂CO₃
LiOH
Ca(OH)₂
Al(OH)₃
4Al(OH)₃
MgSO₄
2MgSO₄
Al₂(SO₄)₃
3Al₂(SO₄)₃

Section 5:

2K₂SO₄
CO₂
2CH₄
4Be(OH)₂
Aluminium iodide
Water
Sodium sulfate

Section 6:

Ga(OH)₃ (s)
Ga(OH)₃ (l)
Ga(OH)₃ (aq)
Cl₂(aq)
3NaCl(s)
3NaCl(aq)
A solution of potassium iodide
Melted barium fluoride