**LeChatelier--Stress--Want--Shift**

 **Changes in Factors** **Effect on the Equilibrium Constant Kc**

**1)** More products added in equal amounts Kc decreases (Equilibrium shifts left)

**2)** More reactants added in equal mole amounts Kc increases (Equilibrium shifts right)

**3)** Increase in pressure Kc increases/decreases depending on the mole ratio reactants to products(Equilibrium shifts right/left)

**4)** Decrease in pressure Kc increases/decreases depending on the mole ratios of reactants to products (Equilibrium shifts right/left)

**5)** Increase in temperature (Exothermic) Kc decreases (Equilibrium shifts left)

**6)** Decrease in temperature (Exothermic) Kc increases (Equilibrium shifts right)

**7)** Increase in temperature (Endothermic) Kc increases (Equilibrium shifts right)

**8)** Decrease in temperature (Endothermic) Kc decreases (Equilibrium shifts left)