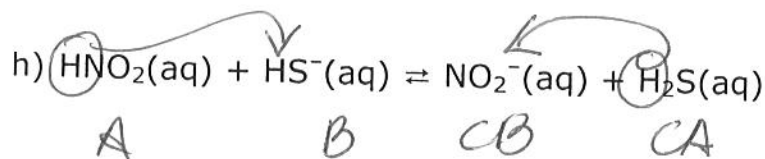
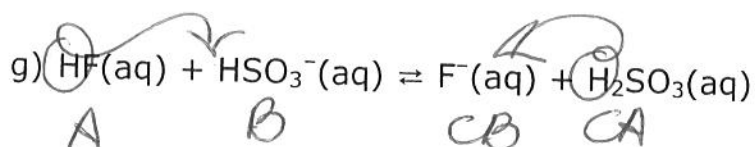
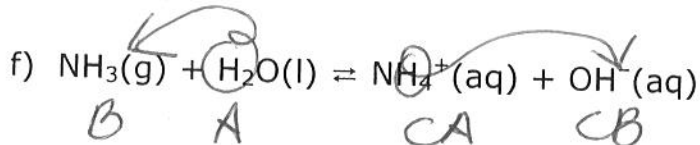
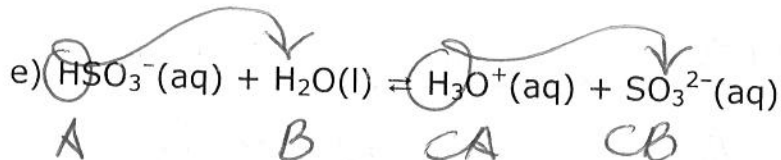
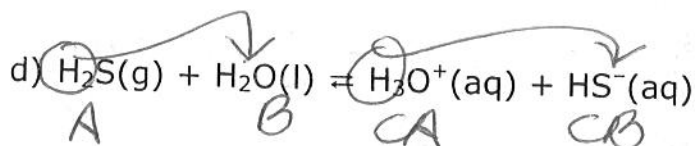
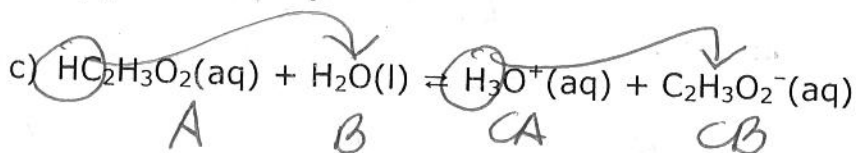
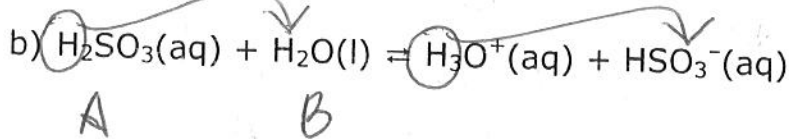
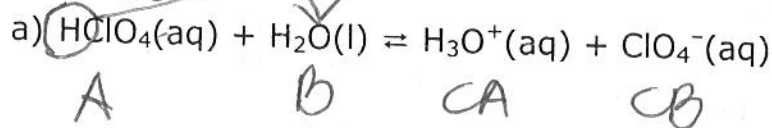


* the larger the K_A the stronger the weak acid is *

Conjugate Pairs Practice Questions

1. Identify the acid, base, conjugate acid and conjugate base for each of the following.



Chemistry 40S

2. Complete the equation for the reaction of each of the following with water. Indicate whether the ion or molecule is an acid or base, and whether each reaction is explained by Arrhenius, Bronsted-Lowry, or both.

$K_A \text{HI} =$
a very
large #

