

Testing the Properties of ABS and dH₂O

Acids = H-NM

Bases = M-OH (for now)

Salts = M-NM

Chemicals:

HCl(aq) hydrochloric acid (only when dissolved in water = aqueous soln)

NaOH(aq) sodium hydroxide base (dissolved in water = aqueous soln)

NaCl(aq) sodium chloride salt (dissolved in water = aqueous soln)

dH₂O(l) distilled H₂O = water with all the minerals i.e. salts removed

	HCl(aq)	NaOH(aq)	NaCl(aq)	dH ₂ O(l)
Test for conductivity	+	+	+	—
Test for Red litmus	—	<i>turns blue</i>	—	—
Test for Blue litmus	<i>turns red</i>	—	—	—
Stupid Cobalt Chloride Paper Test <i>for water</i>	+	+	+	+
Reaction with Mg metal	+ <i>H₂(g)</i>	—	—	—
Reaction with calcium carbonate	+ <i>CO₂(g)</i>	—	—	—

What do acids do?

- conduct elec in water
- turn blue litmus red "BRA"
- react with metals (M) to form $H_2(g)$
- react with carbonates to form $CO_2(g)$

What do bases do?

- conduct elec in water
 - turn red litmus blue
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What do salts do?

- conduct elec in water
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What does distilled water do?

- nada !
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