Answer Key

Formulas and Nomenclature Binary Ionic Compounds Worksheet

Name the following compounds.

1. KCl potassium chloride
2. Li_2O lithium oxide
3. $CaBr_2$ calcium bromide
4. LiH lithium hydride
5. $MgBr_2$ magnesium bromide
6. K_2O potassium oxide
7. ZnO zinc oxide
8. SrS strontium sulfide
9. CaS calcium sulfide
10. Ag_2S silver sulfide
11. ZnF_2 zinc fluoride
12. Ag_3N silver nitride
13. NaF sodium fluoride
14. BaO barium oxide
15. Na_2S sodium sulfide
16. $AlBr_3$ aluminum bromide
17. Li_3N lithium nitride
18. KF potassium fluoride
19. SrI_2 strontium iodide
20. MgO magnesium oxide
21. Al_2O_3 aluminum oxide
22. CaH_2 calcium hydride

Write formulas.

1. sodium bromide NaBr
2. calcium oxide CaO
3. silver chloride AgCl
4. silver oxide Ag_2O
5. aluminum nitride AlN
6. zinc iodide ZnI_2
7. magnesium nitride Mg_3N_2
8. calcium hydride CaH_2
9. potassium phosphide K_3P
10. calcium fluoride CaF_2
11. sodium nitride Na_3N
12. magnesium chloride $MgCl_2$
13. calcium chloride $CaCl_2$
14. potassium iodide KI
15. aluminum chloride $AlCl_3$
16. barium chloride $BaCl_2$
17. sodium chloride NaCl
18. silver bromide AgBr
19. magnesium hydride MgH_2
20. zinc chloride $ZnCl_2$
21. zinc sulfide ZnS

Formulas and Nomenclature Binary Molecular Worksheet

Name the following compounds.

1. N_2O_4 dinitrogen tetroxide
2. N_2O dinitrogen monoxide
3. P_2O_5 diphosphorus pentoxide
4. Cl_2O_7 dichlorine heptoxide
5. CO_2 carbon dioxide
6. OF_2 oxygen difluoride

Write formulas.

1. phosphorus pentachloride PCl_5
2. carbon monoxide CO
3. carbon tetrachloride CCl_4
4. nitrogen trifluoride NF_3
5. sulfur hexafluoride SF_6
6. dinitrogen trioxide N_2O_3

Naming Ionic Compounds Worksheet One

Give the name of the following ionic compounds:

1) Na_2CO_3 sodium carbonate
2) NaOH sodium hydroxide
3) $MgBr_2$ magnesium bromide
4) KCl potassium chloride
5) $FeCl_2$ iron (II) chloride
6) $FeCl_3$ iron (III) chloride
7) $Zn(OH)_2$ zinc hydroxide
8) $BeSO_4$ beryllium sulfate
9) CrF_2 chromium (II) fluoride
10) Al_2S_3 aluminum sulfide
11) PbO lead (II) oxide
12) Li_3PO_4 lithium phosphate
13) TiI_4 titanium (IV) iodide
14) Co_3N_2 cobalt (II) nitride
15) Mg_3P_2 magnesium phosphide
16) $Ga(NO_2)_3$ gallium nitrite
17) Ag_2SO_3 silver sulfite
18) NH_4OH ammonium hydroxide
19) $Al(CN)_3$ aluminum cyanide
20) $Be(CH_3COO)_2$ beryllium acetate

NO_3^- = nitrate
NO_2^- = nitrite

For the following compounds, give the formulas

22) sodium phosphide Na_3P

23) magnesium nitrate $Mg(NO_3)_2$

24) lead (II) sulfite $PbSO_3$

25) calcium phosphate $Ca_3(PO_4)_2$

26) ammonium sulfate $(NH_4)_2SO_4$

27) silver cyanide $AgCN$

28) aluminum sulfide Al_2S_3

29) beryllium chloride $BeCl_2$

30) copper (I) arsenide Cu_3As

31) iron (III) oxide Fe_2O_3

32) gallium nitride GaN

33) iron (II) bromide $FeBr_2$

34) vanadium (V) phosphate $V_3(PO_4)_5$

35) calcium oxide CaO

36) magnesium acetate $Mg(CH_3COO)_2$

37) aluminum sulfate $Al_2(SO_4)_3$

38) copper (I) carbonate Cu_2CO_3

39) barium oxide BaO

40) ammonium sulfite $(NH_4)_2SO_3$

41) silver bromide $AgBr$

42) lead (IV) nitrite $Pb(NO_2)_4$