**2 Good Morning Practice Questions for You! GO!!!**

**Find the net ionic equation:**

BCE 2 NaOH (aq) + H2SO4 (aq) --> Na2SO4(aq) + 2 H2O (l)

I.E. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

N.I.E. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Potassium chlorate decomposes upon slight heating in the presence of a catalyst according to the reaction below:

**2 KClO3(s) → 2 KCl(s) + 3O2(g)**

In a certain experiment, 40.0 g KClO3 is heated until it completely decomposes.

i) Where is the catalyst written in the equation?

ii) Where would the energy be written in the equation?

iii) What is the theoretical yield of oxygen gas?

iv) The experiment is performed and the oxygen gas is collected and its mass is found to be 14.9 g.

What is the percent yield for the reaction?